

Directions: Read Section 10.8 and answer the following:

1. When a liquid evaporates, is this an endothermic or an exothermic process?
2. What is the "heat of vaporization"?
3. Why does water have an unusually large heat of vaporization?
4. What is the process by which vapor molecules become a liquid?
5. What is the definition of "vapor pressure of a liquid"?
6. How does one measure vapor pressure?
7. Why is the vapor pressure also called the "equilibrium vapor pressure"?
8. How does a liquid's vapor pressure change with temperature?
9. If something is volatile, does it have a high or a low vapor pressure? Why? Explain in terms of temperature and kinetic energy.
10. What is the "enthalpy of fusion"?
11. a) What is the difference between the "melting point" and the "normal melting point"?

b) What is the difference between the "boiling point" and the "normal boiling point"?
12. Draw the heating curve for water on the back of this worksheet and label where the following would be found: melting point, boiling point, heat of fusion, heat of vaporization.