

Worksheet # C39: Solubilities and Double Displacement

For each of the following double displacement reactions, first write the complete balanced equation. Leave room for the phases (s) or (aq). Then use the solubility rules to determine if each of the products will be (aq) or (s). Finally, answer the question, "will this reaction occur?" The reaction will only occur if one of the products is a precipitate (a solid), a liquid, or a gas.

Will this
reaction occur?

1. Potassium phosphate + zinc nitrate

2. Zinc chloride + strontium nitrate

3. Calcium hydroxide + iron(III) iodide

4. Barium bromide + silver (I) acetate

5. Aluminum sulfate + magnesium bromide

6. Tin(II) bromide + ammonium phosphate

7. Aluminum iodide + silver nitrate

8. Sodium chloride + lead(II) acetate
