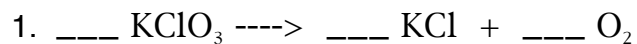


## Worksheet # C46: Grams to Grams Using Balanced Equations

Directions: Balance each of the following equations and answer the questions. *Make sure you show your set-up.* For each question, getting the set-up right is almost as important than getting the answer right. (Remember sig figs, too...)



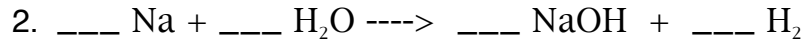
a. What is the molecular mass of :

i.  $\text{KCl}$ ? \_\_\_\_\_

ii.  $\text{KClO}_3$ ? \_\_\_\_\_

b. How many grams of  $\text{KCl}$  could you get if you decomposed 25.8 grams of  $\text{KClO}_3$ ?  
\_\_\_\_\_

c. How many grams of  $\text{KClO}_3$  would you need to decompose in order to make 24.61 grams of  $\text{O}_2$ ?  
\_\_\_\_\_



a. What is the molecular (or atomic) mass of :

i. Na? \_\_\_\_\_

ii. NaOH? \_\_\_\_\_

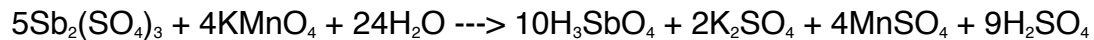
b. How many grams of pure sodium metal would it take to produce 150. grams of sodium hydroxide?

\_\_\_\_\_

c. How many grams of  $\text{H}_2$  would you make in the process of making 150. grams of NaOH?

\_\_\_\_\_

3. Here's this wild equation again:



a. What is the molecular mass of :

i.  $\text{Sb}_2(\text{SO}_4)_3$ ? \_\_\_\_\_

ii.  $\text{MnSO}_4$  ? \_\_\_\_\_

a. How many grams of  $\text{Sb}_2(\text{SO}_4)_3$  would it take to produce 250.0 grams of  $\text{MnSO}_4$  ?

\_\_\_\_\_