

## Fake Quiz For the Chemical Bonding Test

I. Match the terms on the right with the phrases on the left (terms may be used more than once):

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|---|---|
| <ol style="list-style-type: none"> <li>1. _____ Phosphorus has _____ valence electrons.</li> <li>2. _____ Gallium has _____ valence electrons.</li> <li>3. _____ Atom that wants 8 valence electrons when bonding.</li> <li>4. _____ Magnesium has _____ valence electrons.</li> <li>5. _____ Bond that involves the transfer of electrons.</li> <li>6. _____ Shape of carbon dioxide.</li> <li>7. _____ Shape of water.</li> <li>8. _____ Sulfur has _____ valence electrons.</li> <li>9. _____ Atom that wants 6 valence electrons when bonding.</li> <li>10. _____ Bond between metals.</li> <li>11. _____ Kind of electrons involved in bonds.</li> <li>12. _____ Shape of methane (CH<sub>4</sub>).</li> <li>13. _____ Atom that wants 4 valence electrons when bonding.</li> <li>14. _____ Bond between a metal and a non-metal.</li> </ol> | <ol style="list-style-type: none"> <li>a. core</li> <li>b. valence</li> <li>c. linear</li> <li>d. trigonal planar</li> <li>e. trigonal pyramidal</li> <li>f. tetrahedral</li> <li>g. bent</li> <li>h. ionic</li> <li>m. non-polar covalent</li> <li>n. polar covalent</li> <li>o. metallic</li> <li>p. electronegativity</li> <li>q. chlorine</li> <li>r. hydrogen</li> <li>s. beryllium</li> <li>t. boron</li> <li>u. one</li> <li>v. two</li> <li>w. three</li> <li>x. four</li> <li>y. five</li> <li>z. six</li> <li>aa. seven</li> <li>bb. eight</li> </ol> |
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II. Right now you're breathing N<sub>2</sub>(g), O<sub>2</sub>(g), Ar(g), CO<sub>2</sub>(g), and H<sub>2</sub>O(g) (along with a few other trace gases). Draw the Lewis dot structures for each of these in the boxes below:

N <sub>2</sub>	Ar	
O <sub>2</sub>	CO <sub>2</sub>	

Draw the Lewis Structure and 3D Structural Drawing for each of these:

$\text{BCl}_3$		
$\text{AsCl}_3$		
$\text{SiS}_2$		
$\text{OSe}_2$		

How well prepared are you for this test right now? \_\_\_\_\_