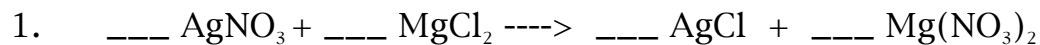


## Limiting Reactant Practice Problems

Directions: For each of the following, balance the equation and work out the problems. Remember to show your set-ups in each case and always pay attention to those sig figs.

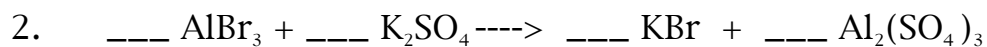


If you have 5.00 grams of silver nitrate and 5.00 grams of magnesium chloride....

a. Which is the limiting reactant? \_\_\_\_\_

b. Which is the excess reactant? \_\_\_\_\_

c. How many grams of silver chloride would you be able to make?

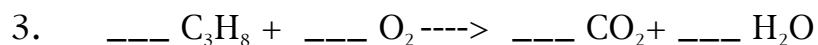


If you have .250 grams of aluminum bromide and .350 grams of potassium sulfate....

a. Which is the limiting reactant? \_\_\_\_\_

b. Which is the excess reactant? \_\_\_\_\_

c. How many grams of aluminum sulfate would you be able to make?

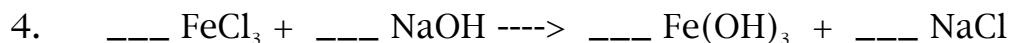


If you have 100. grams of  $\text{C}_3\text{H}_8$  and 200. grams of  $\text{O}_2$ ....

a. Which is the limiting reactant? \_\_\_\_\_

b. Which is the excess reactant? \_\_\_\_\_

c. How many grams of  $\text{CO}_2$  would you be able to make?



If you have 46.9 grams of iron III chloride and 18.5 grams of sodium hydroxide....

a. Which is the limiting reactant? \_\_\_\_\_

b. Which is the excess reactant? \_\_\_\_\_

c. How many grams of iron III hydroxide would you be able to make?