

3.. For each of the following, first write the balanced equation (be). Then write out (cie) the complete ionic equation and (nie) the net ionic equation for each.

a) sodium phosphate and calcium acetate

be) _____

cie) _____

nie) _____

b) barium chloride and sodium hydroxide

be) _____

cie) _____

nie) _____

4. For these last three, first write the balanced equation (be). Then write out just (nie) the net ionic equation for each.

a.) copper metal in silver nitrate

be) _____

nie) _____

b.) lithium sulfide and aluminum bromide

be) _____

nie) _____

c.) ammonium carbonate and magnesium sulfate

be) _____

nie) _____