



### Calculating The Percents (Two Equations, Two Unknowns)

5. Neon has two isotopes, Ne-22 and Ne-20. The approximate mass of Ne-22 is 22.000 amu and the approximate mass of Ne-20 is 20.000 amu. The atomic mass of neon is 20.1798 amu. Determine the proportions of Ne-22 and Ne-20 in a naturally occurring sample of neon.

6. Boron has two isotopes, B-10 and B-11. B-10 has an approximate mass of 10.000 amu, and B-11 has an approximate mass of 11.000 amu. The atomic mass of boron is 10.811 amu. Determine the proportions of B-10 and B-11 in a naturally occurring sample of boron.

7. Europium has two isotopes, Eu-151 and Eu-152. Eu-151 has an approximate mass of 151.000 amu and Eu-152 has an approximate mass of 152.000 amu. The atomic mass of europium is 151.964 amu. Determine the proportions of Eu-151 and Eu-152 in a naturally occurring sample of europium.