

ZnCl_{???} Lab Follow Up

1. Using two sig figs for the ???, what was your lab value for ZnCl_{???}? _____
2. The correct value for the ??? was supposed to be 2.0. Calculate the following, showing your set ups and circling your answers.
 - a. Percent yield
 - b. Percent error:
3. Just to make sure you're comfortable working with this kind of data (there will be a question exactly like this on the next test), calculate the value of the ??? for this lab:

In a lab just like the one we just did, an empty beaker was weighed. A small sample of pure manganese powder was added to the beaker, and about 5 mL of 6M HCl was added. After the fizzing stopped, the beaker was put on a hot plate, and all the excess water and HCl was boiled away. All that was left was a white powder in the beaker. The beaker with the powder was weighed. The data recorded was:

Empty Beaker	35.892 g
Manganese (before)	.170 g
MnCl _{???} (white powder) + beaker	36.613 g

- a. What was the value of the ??? in MnCl_{???} ? (Show all set-ups.)

- b. If the correct answer for the ??? was 5.00, what was the percent yield for this lab?